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Title: Chapter 25 Nuclear Chemistry 1 Chapter 25 Nuclear Chemistry 2 Chemical vs. Nuclear Reactions Chemical Reactions- A rearrangement of atoms and molecules by breaking and forming bonds -involves electrons Nuclear Reactions-Combining, splitting or decay the nuclei of atoms. 3 Nuclear Chemistry involves the NUCLEUS of the atom.

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Chapter 25 Nuclear Chemistry . Radioactivity •Roentgen was studying fluorescence when he accidentally discovered the X-ray •Becquerel then discovered the radioactivity •Marie Curie then continued the research to discover more radioactive elements . Terms you need to know

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Chapter 25 Nuclear Chemistry

Chapter 25: Nuclear Chemistry. 25.1: Nuclear Radiation. Marie Curie (1867-1934) and Pierre Curie (1859-1906) were able to show that rays emitted by uranium atoms caused fogging in photographic plates. Slideshow 5980037 by tanisha-wilkinson

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Chapter 25: Nuclear Chemistry. Explain how an unstable nucleus releases energy. Compare and contrast the three main types of nuclear radiation. Identify the two ways transmutation can occur. Answer...

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CHAPTER 25. Nuclear Chemistry: Radiation, Radioactivity & its Applications. Alta Chemistry. Nuclear Chemistry. Alta Chemistry. Nuclear Chemistry deals with changes in the nucleus. The nucleus of an atom contains. Protons - Positively Charged.

Nuclear Chemistry: Radiation, Radioactivity & its Applications

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Section 25.1 Nuclear Radiation799 Marie Curie was a Polish scientist whose research led to many discoveries about radiation and radioactive elements. In 1903 she and her husband Pierre, along with Antoine Henri Becquerel, won the Nobel Prize in physics for their work on radioactivity.

25.1 Nuclear Radiation 25

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806 Chapter 25 Nuclear Chemistry Figure 25-2 Both Pierre and Marie Curie played important roles in founding the field of nuclear chemistry Marie Curie went on to show that unlike chemical reactions, radioactivity is not affected by changes in physical conditions such as temperature and pressure She is the only person in history to

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 $BE = 0.1587 \text{ amu}$
 $1 \text{ amu} = 1.49 \times 10^{-10} \text{ J}$
 $BE = 2.37 \times 10^{-11} \text{ J} = 1.25 \times 10^{-12} \text{ J}$
 23.2 ...

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Nuclear Chemistry Nuclear reactions are reactions that affect the nucleus of the atom. 5. Radioactivity is the phenomenon of radiation (particles and/or energy) being ejected spontaneously by an unstable nucleus until it reaches a more stable arrangement. 6. Nuclear Stability is determined by the ratio of

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protons to neutrons in the nucleus. 7.

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Chapter 12 - Physical Properties of Solutions Chapter 13 - Chemical Kinetics Chapter 14 - Chemical Equilibrium Chapter 15 - Acids and Bases Chapter 16 - Acid-Base Equilibria and Solubility Equilibria Chapter 17 - Chemistry in the Atmosphere Chapter 18 - Entropy, Free Energy, and Equilibrium Chapter 19 - Electrochemistry Chapter 20 - Metallurgy and the Chemistry of Metals Chapter 21 ...

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692 Chapter 16 Nuclear Chemistry 16.1 The Nucleus and Radioactivity Our journey into the center of the atom begins with a brief review. You learned in Chapter 3 that the protons and neutrons in each atom are found in a tiny, central nucleus that measures about 1/100,000 the diameter of the atom itself. You also learned

Chapter 16 Nuclear Chemistry

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