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Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key

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Balancing

Balancing Chemical

Equations Answer Key

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Balance the equations

below: 1) $1 \text{ N}_2 + 3 \text{ H}_2$

$\rightarrow 2 \text{ NH}_3$ 2) 2 KClO_3

$\rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) 2

$\text{NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF}$

$+ 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2$

$\rightarrow 2 \text{ H}_2\text{O}$ 5) 1 Pb(OH)

$2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} +$

1 PbCl_2 6) $2 \text{ AlBr}_3 + 3$

$\text{K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}$

$2(\text{SO}_4)_3$ 7) $1 \text{ CH}_4 + 2$

$\text{O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow$

$3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 9) 2 C

$8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16$

$\text{CO}_2 + 18 \text{ H}_2\text{O}$...

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Chapter 7
Worksheet #1
Balancing Chemical
Equations

Chapter 7 Worksheet
#1 Balancing Chemical
Equations

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Equations ...

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Balancing Chemical

Equations – Answer

Key Balance the

equations below: 1) 1

$\text{N}_2 + 3 \text{H}_2 \rightleftharpoons 2 \text{NH}_3$ 2)

$2 \text{KClO}_3 \rightleftharpoons 2 \text{KCl} + 3$

O_2 3) $2 \text{NaCl} + 1 \text{F}_2 \rightleftharpoons$

$2 \text{NaF} + 1 \text{Cl}_2$ 4) 2H_2

$+ 1 \text{O}_2 \rightleftharpoons 2 \text{H}_2\text{O}$ 5) 1

$\text{Pb}(\text{OH})_2 + 2 \text{HCl} \rightleftharpoons 2$

$\text{H}_2\text{O} + 1 \text{PbCl}_2$ 6) 2

$\text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightleftharpoons 6$

$\text{KBr} + 1 \text{Al}_2(\text{SO}_4)_3$ 7) 1

$\text{CH}_4 + 2 \text{O}_2 \rightleftharpoons 1 \text{CO}_2 +$

$2 \text{H}_2\text{O}$ 8) $1 \text{C}_3\text{H}_8 + 5$

$\text{O}_2 \rightleftharpoons 3 \text{CO}_2 + 4 \text{H}_2\text{O}$

9)

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Chapter Chloride

Chlorine. Chem1211

Chapter 3 ...

Chapter 7

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Equations Key ...

Question: Chapter 7

Worksheet #1 Balance

The Equations Below:

4) $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + \text{H}_2$

5) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

7) $\text{H}_2 + \text{Cl}_2 \rightarrow 2 \text{HCl}$

8) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

12) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$

17) $2 \text{HCl} + \text{Ca} \rightarrow \text{CaCl}_2 + \text{H}_2$

18) This

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problem has been solved! See the answer. Show transcribed image text. Expert Answer . 1. $H_2 + N_2 \rightarrow NH_3$ Notice that there are 2 N on the left side but only 1 on the right so let's fix that.

Solved: Chapter 7 Worksheet #1 Balance The Equations Below ...

Chapter 7:
Stoichiometry of

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Chemical Reactions.

Search for: 7.1: Writing
and Balancing
Chemical Equations.

Learning Objectives. By
the end of this section,
you will be able to:

Derive chemical
equations from
narrative descriptions
of chemical reactions.

7.1: Writing and Balancing Chemical Equations | General

...

Title: Balancing
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Balancing Equations

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Equations Worksheet 1

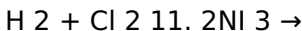
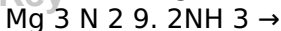
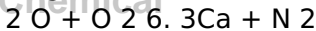
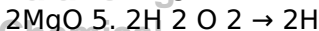
- ANSWERS 1. $2\text{H}_2 + \text{O}_2$

$\rightarrow 2\text{H}_2\text{O}$ 2. $2\text{Na} + \text{Cl}_2$

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Balancing Chemical Equations

Worksheet 1

If you also get perplexed in balancing chemical equations,

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follow the tips for correct balancing chemical equations worksheet answers. Tip # 1: When you are trying to balance the chemical equations, you should remember that you can only change the value of coefficient in front of the element or compound, and not the subscript.

49 Balancing Chemical Equations

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Worksheets [with Answers]

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Equations - Before you

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notices you have to

read the script.

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be helpful after you are

trying to lose or

maintain. As soon as

you're trying to lose or

maintain your weight,

they may be helpful.

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NH_3 _ 1) _ N_2 + _ H_2 2)

2

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equations below: 1) 1

$\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$ 2) 2

$\text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ 3)

$2 \text{NaCl} + 1 \text{F}_2 \rightarrow 2 \text{NaF} +$

1Cl_2 4) $2 \text{H}_2 + 1 \text{O}_2$

$\rightarrow 2 \text{H}_2\text{O}$ 5) $1 \text{Pb}(\text{OH})_2 +$

$2 \text{HCl} \rightarrow 2 \text{H}_2\text{O} + 1 \text{PbCl}_2$

6) $2 \text{AlBr}_3 + 3 \text{K}_2\text{SO}_4$

$\rightarrow 6 \text{KBr} + 1 \text{Al}_2(\text{SO}_4)_3$

7) $1 \text{CH}_4 + 2 \text{O}_2 \rightarrow 1$

$\text{CO}_2 + 2 \text{H}_2\text{O}$ 8) $1 \text{C}_3\text{H}_8$

$+ 5 \text{O}_2 \rightarrow 3 \text{CO}_2 +$

$4 \text{H}_2\text{O}$ 9) $2 \text{C}_8\text{H}_{18} +$

$25 \text{O}_2 \rightarrow 16 \text{CO}_2 + 18 \text{H}_2\text{O}$

10) $1 \text{FeCl}_3 + 3 \dots$

H₂SO₄ NaNO₃ HNO₃ Na

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to automate your
business.

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