Chemactivity 9 Shell Model Answers

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4.The IE of Li+ should be larger than the IE of He because both atoms have 2 electrons in the 1st shell and Li+ has a core charge of +3 whereas He only has a core charge of +2. ... ChemActivity 9. 1.a) Two. b) Lower energy peak (1. s ... In the model, the C–C single bond energy is 376 kJ/mole and the C–C double bond energy is 720 kJ/mole. ...

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ChemActivity 9 1. a) Two. b) Lower energy peak (1s) is 2 x the intensity of the higher energy peak (2s). c) The nuclear charge for H, He, and Li is 1, 2, and 3, respectively. Therefore, the

electrons in the first shell will be held most tightly by Li and least tightly by H. d) H and Li have the same core charge; the electron is farther away in Li.

Answers to Exercises and Problems

Solved: Meads ChemActivity 9 Is Electron Arrangement (Wher ... 9.Calculations:a) A(0), A(0)b) C(0), C(0) Assume that A, B, and C have seven valence electrons and that there is a single bond in all molecules. c) A(-0.50), B(+0.50) d) A(–0.10), C(0.10) ChemActivity 22 Chem;GI;Answers

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ChemActivity 4 The Shell Model (I) 17 Model 2: Simple Model Diagrams for Hydrogen and Helium Atoms. One simple model of the hydrogen atom pictures the H atom as a nucleus of charge +1 surrounded by an electron at some distance, as shown in Figure 1. Figure 1. Model diagram of a hydrogen atom. +1 electron the electron "sees" a

The Shell Model (I) - Science with Mr. Louie

9.a) Diagram (a) should have a nuclear charge of +11, two electrons in the first circle (shell), and nine electrons in the second circle (shell). Diagram (b) should have a core charge of +9 and a circle with nine electrons (at about the same distance as the second circle in diagram (a).

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ChemActivity 3 Coulombic Potential Energy 9 5. Recall that a 1H atom consists of a proton as the nucleus and an electron outside of the nucleus. Is the potential energy, V, of a hydrogen atom a positive or negative number? Explain your answer. Model 2: Ionization Energy.

Chemistry: A Guided Inquiry 5e
Given the shell model of the atom, why do you think that Lewis proposed a maximum of two electrons for hydrogen and eight for carbon, nitrogen, oxygen, and fluorine atoms? 18. Answer the following for the phosphorus atom: a. How many valence electrons does have? b. What is the Lewis representation for P? c.

Solved: ChemActivity 10 Covalent And Lonic Bonds (Why Do A ...
Based on your answer to CTQ 13 and the ionization energy data, Table 1 of ChemActivity 4: Shell Model (I), is the radius of the valence shell of Na larger, smaller or the same as the radius of the valence shell of Li? 15. Consider the models of Ne and Na shown in Models 4 and 5. Explain how the core

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Updates to Bohr Model:-Electrons are NOT in circular orbits around nucleus.-Electrons are in a 3-D region around the nucleus called atomic orbitals.-The atomic orbital describes the probable location of the electrion Quantum Mechanical Model of the Atom.

Quantum Mechanical Model of the Atom

ChemActivity 8 Photoelectron Spectroscopy ... a shell model of the atom, and noted that the number of valence electrons in the outermost shell is related to the position of the element in the periodic table, and therefore is an important factor in determining the physical and chemical properties of the element. ... In order to answer this ...

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Chemactivity 8 Answers ChemActivity 8. 1. 140.3 MJ/mole. 2.a) 3. ChemActivity 9. 1.a) Two. b) Lower energy peak (1. s) is 2 x the intensity of the higher energy peak (2. s). c) The nuclear charge for H, He, and Li is 1, 2, and 3, respectively.