

Chemical Quantities Chapter 10 Quiz

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the SI unit representing 6.02×10^{23} representative particles of a substance: Avogadro's number: 6.02×10^{23} particles: standard temperature and pressure (00 C, 1 atm) the temperature and pressure at which one mole of gas occupies a volume of 22.4 L: molar volume: volume of a gas that contains one mole of the gas, is 22.4 L at STP: Avogadro ...

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find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole. atom $6.02 \cdot 10^{23}$ O 2 $6.02 \cdot 10^{23}$ ion ...

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Chapter 10 Review: Chemical Quantities & Calculations (aka - mole & volume calculations, % composition, empirical & molecular formulas) For the test, be sure that you are familiar, and able to complete, the problems below as well as the problems found on the following worksheets: "Molecular Weight and Mole Calculations," "Percent Comp. & Chemical

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