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### Chemistry Dimensions 1

Sol. Contribution by atoms Q present at the eight corners of the cube =  $\left(\frac{1}{8}\right) \times 8 = 1$  Contribution by atom P present at the body centre = 1 Thus, P and Q are present in the ratio 1:1.

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$= 100 \text{ g} \cdot 1.41 \text{ g mL}^{-1} \cdot \frac{100 \text{ g}}{1.41 \text{ g mL}^{-1}} \cdot 1.41 \text{ g mL}^{-1} \cdot 100 \text{ g} = 70.92 \text{ mL} = 70.92 \times 10^{-3} \text{ L} = 70.92 \times 10^{-3} \text{ L}$

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## NCERT Solutions for Class 11 Chemistry Chapter 1

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Rajasthan Board RBSE Class 12 Chemistry Chapter 1 Solid State RBSE Class 12 Chemistry Chapter 1 Text Book Questions RBSE Class 12 Chemistry Chapter 1 Multiple Choice Questions. Question 1. The number of atoms in body centred cubic arrangement is-(a) 1 (b) 2 (c) 4 (d) 6

## RBSE Solutions for Class 12 Chemistry Chapter 1 Solid State

Solubility is the property of a solid, liquid or gaseous chemical substance called solute to dissolve in a solid, liquid or gaseous solvent. The solubility of a substance fundamentally depends on the physical and chemical properties of the solute and solvent as well as on temperature, pressure and presence of other chemicals (including changes to the pH) of the solution.

## Solubility - Wikipedia

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An alloy is a solid solution consisting of a metal (like iron) with some other metals or nonmetals dissolved in it. Steel, an alloy of iron and carbon and small amounts of other metals, is an example of a solid solution. Table  $\{\text{PageIndex}\{1\}\}$  lists some common types of solutions, with examples of each.

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