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the periodic table - an  
elements mass listed  
on the table equals the  
mass of one mole of  
that element where  
can you find the molar  
mass of an element?  
the periodic table - find  
the mass of each  
element in a molecule  
multiplied by the

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number of times that  
element appears in the  
molecular formula.

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## **Atoms and Moles**

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atomic mass. the mass of an atom expressed in atomic mass units (AMU) mole. the standard unit used to measure the amount of a substance in a chemical reaction.

molar mass. the mass in grams of 1 mole of a substance. avogadro's number.  $6.022 \times 10^{23}$ , it is the number of particles of atoms or



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molecules in 1 mole.

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The mole is the most  
common unit used to  
express the quantity of  
a chemical substance.  
For all solids, liquids,

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and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at  $0^{\circ}\text{C}$  and 1 atmosphere pressure (known as standard temperature and pressure or STP ), one mole of any gas occupies approximately 22.4 liters.

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atomic mass is the mass of 1 atom of a particular element and molar mass is the mass of 1 mole of an atom or

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molecule how many  
atoms in 1 mole of  
water?  $6.022 \times 10^{23} \times 3$

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chemistry i (tesc 141)  
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stoichiometry mole-a  
unit of measurement  
that expresses the  
amount of atoms,  
molecules or some  
other unit The number  
of items in one mole is

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commonly referred to  
Avogadro's number  
which equals

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Stoichiometry which  
generally used for the  
conversion of number  
of moles in mol of any  
given chemical species  
to the mass of that  
species in grams  
through molar mass is  
commonly known as  
mass.

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mass of one mole of a  
pure substance.  
Avogadro's number. ...  
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terms.

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Mole: Mole is the SI unit for measuring the amount of substance. One mole can be defined as the amount of substance that contains Avogadro number of . atoms, molecules, ions or

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electrons.

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required to be  
produced ... -  
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It is the number of atoms present is 12 amu of C-12. Where one mole of a substance contains an Avogadro's Number of particles. 
$$1 \sim \text{mole} = 6.022 \times 10^{23} \sim \text{particles}$$

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Which has the smaller  
mass, 1 mole of  
 $\text{He}$  atoms or  
4 moles of  $\text{H}$   
atoms?

Avogadro's Number:

The quantity that  
relates the amount of  
particles to moles is  
known as Avogadro's

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**Which has the  
smaller mass, 1  
mole of He atoms or  
4 moles ...**

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