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mass and molecular formulas.

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10 Mole. STUDY. PLAY. mole. the SI base
unit used to measure the amount of a
substance. mole. is defined as the
number of carbon atoms in exactly 12g

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of pure carbon- 12. Avogadro's number.
the number 6.02×10^{23} in honor of the
Italian physicist and lawyer Amedeo
Avogadro.

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slides. • Explain how a mole is

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Measuring Matter 1. pair 2. 5 3. dozen 4.
gross 5. 200 6. ream 7. 6,000,000,000 8.
0.5 mol 9. 6.02 $\times 10^{23}$ 10. four moles 11.
6.02 $\times 10^{23}$ Cu atoms 12. 4 $\times 10^{23}$

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1 mol CH₆ 0.02 10 molecules CH₁₃. 23 1
mol Xe 6.02 10 molecules Xe 14. 23 2 2

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10 161 Section 10.1 Measuring Matter
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323–324 1. Zinc (Zn) is used to form a

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corrosion-inhibiting surface on
galvanized steel. Determine the number
of Zn atoms in 2.50 mol of Zn. 2.50 mol
Zn

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State. thermal contraction. thermal
expansion. freezing point. boiling point.

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a decrease in a material's volume when the temperature is decr.... An increase in the size of a substance when the temperature is.... The temperature at which a liquid changes into a solid.

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