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Heat Capacity, Molar

Heat Capacity, and

Specific Heat. The heat

capacity, C , is the

amount of heat, q ,

required to raise the

temperature, ΔT , of an

object by $1\text{ }^{\circ}\text{C}$. The

three variables are

related by the equation

$q = C\Delta T$ The value

of C in this equation,

and likewise the

magnitudes of q and

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ΔT , pertain to a certain sample and depend on the amount.

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16-4 1. When a 12.8-g
sample of KCl dissolves
in 75.0 g of water in a
calorimeter, the
temperature drops
from 31.0°C to 21.6°C.
Calculate ΔH for the pr

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process

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This process is called "measuring heat transfer" calorimetry.

Calorimeters are devices used in measuring heat flow. In calorimeters; Heat

Absorbed = Heat Released. ... chemistry ii enthalpy worksheet answers online

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addition to mass
changes, chemical
reactions involve heat
changes associated
with changes in the
substances' internal
energy. Like mass-
based stoichiometry,
these changes are
quantitative.

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3:57:00 PM) Name:

Date: § 16.01a

Thermochemistry:

Calorimetry Most

Important Ideas: 1.

specific heat (s) =

amount of heat

required to raise 1 -g

substance by 1 oC

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2) When 62.3g of a
compound was burned

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in a bomb calorimeter that contained 0.500 L of water the temperature rise of the water in the

calorimeter was 48.00

C. If the heat of combustion of the compound is 1,160 kJ/mol, what is the molar mass of the compound? 3) The

molar heat of combustion of a compound is 1,350 kJ/mol. If 0.875 moles of this

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Lessons. In this lesson,

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Heat Specific heat is

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the amount of heat required to raise one gram of any substance one degree Celsius or Kelvin.

**Thermochemistry
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Worksheet 16 —

Calorimetry

Calorimetry is the

experimental

measurement of heat

(q) produced in

chemical and physical

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processes. Heat can not be measured directly, but temperature changes can be measured. The factor that links these two is heat capacity.

butane.chem.illinois.edu

Name: _____ Hour: _____

Date: _____ Chemistry:

Calorimetry Problems

1. Solve the following problems. As always, include work and show the units to ensure full

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credit. 1. A 445 g sample of ice at -58°C is heated until its temperature reaches -29°C . Find the change in heat content of the system. 2.

Calorimetry

Problems 1 - FREE

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