

## Molarity By Dilution Chemistry If8766 Answers

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### **Molarity By Dilution Chemistry If8766**

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Molarity, or molar concentration, represents the concentration of a solute in a solution. The unit usually used for molarity in chemistry is mol/L and is Page 6/10

### **Chemistry Molarity Worksheet Answers**

Chemistry If8766 . MOLARITY BY DILUTION Name Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid =  $M \times V$  then,

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where the subscripts "1" and "2" refer to the solution before and after the dilution, respectively. Since the dilution process does not change the amount of solute in the solution,  $n_1 = n_2$ . Thus, these two equations may be set equal to one another:

$[M_1V_1 = M_2V_2]$  This relation is commonly referred to as the dilution equation.

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[Books] Molarity By Dilution Chemistry Pg 69 Answer molarity of BaBr<sub>2</sub> solution:  $0.058375 \text{ mol} / 0.165 \text{ L} = 0.35 \text{ M}$  Problem #9: 1.00 L of a solution is prepared by dissolving 125.6 Molarity By Dilution 69 Answers - Page 7/29 Molarity By

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BaBr<sub>2</sub> solution:  $0.058375 \text{ mol} / 0.165 \text{ L} = 0.35 \text{ M}$  Problem #9: 1.00 L of a solution is prepared by dissolving 125.6 g of BaBr<sub>2</sub> in 1.00 L of a solution. Calculate the molarity of the solution.

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Chemistry If8766 Molarity Answers [Books] Molarity By Dilution Chemistry Pg 69 Answer molarity of BaBr<sub>2</sub> solution:  $0.058375 \text{ mol} / 0.165 \text{ L} = 0.35 \text{ M}$  Problem #9: 1.00 L of a solution is

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prepared by dissolving 125.6 g of NaF in it.

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